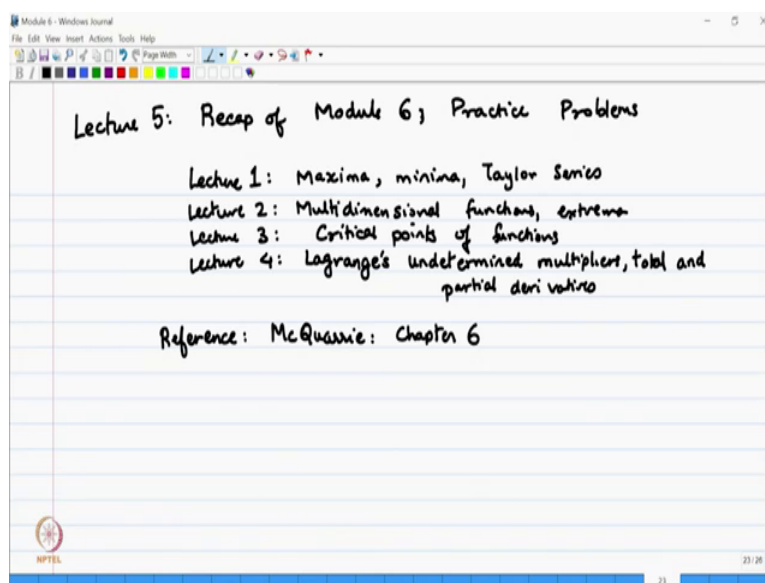


Advanced Mathematical Methods for Chemistry
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Module - 06
Lecture – 05
Recap of Module 6, Practice Problems

So, this will be the last lecture of Module 6 and today I will start by recapping what you have learnt and then, we will do some practice problems. So, to recap in this module, we basically focused on maximization and minimization of functions and we looked at Lagrange's methods of undetermined multipliers which allows us to find maxima or minima of functions subject to certain constraints.

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So, in the first lecture, we learnt about maxima minima and we also learnt about Taylor series. In the second lecture, we did maximization and minimization of multi-dimensional functions. We also looked at Taylor series for multi-dimensional functions and we looked at all the various extreme that you can have. In the third lecture, we classified all the critical points of functions into maxima minima saddle points etcetera, for multi functions of many variables and in the fourth lecture, I talked about Lagrange's undetermined multipliers to find maxima and minima subject to certain constraints. We also talked about total and partial derivatives, how these are related. Now, all of this is

very nicely discussed in Mcquarrie chapter 6. Again chapter 6 has lot of other topics, but all these topics are fairly well covered here.

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Module 6 - Windows Journal

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Problem 1: Taylor Series, maxima and minima

Write the Morse Potential

$$V(r) = D_e \left[1 - e^{-a(r-r_e)} \right]^2 - 1$$

as a Taylor Series in r about $r=r_e$ (1st 3 terms)

Find Extrema of $V(r)$ for $r > 0$

Solution: $V(r_e) = -D_e$; $V'(r_e) = 2D_e(1 - e^{-a(r-r_e)}) a e^{-a(r-r_e)} \Big|_{r=r_e}$

$$= 0$$

$$V''(r_e) = 2aD_e \left[-a e^{-a(r-r_e)} + 2a^2 e^{-2a(r-r_e)} \right] \Big|_{r=r_e}$$

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Now, let us look at some applications of these through some practice problems. So, the first problem that I am going to talk about has to do with Taylor series. Maxima and minima, the question is write the morse potential V is a function of r . V r is a scalar, V is a scalar. It is given by a constant demultiplied by this function. So, 1 minus e raise to minus a r minus r_e , this whole thing is squared and then, you have a minus 1. Write this as a Taylor series in r about r equal to r_e and I am asking you to write the first three terms and find the extrema of V of r for r greater than 0.

So, let us write a solution. So, suppose I take V of r_e is equal to, if I put r equal to r_e , then this term. So, you have e raise to e^2 . So, this exponent goes to 0. So, exponent of 0 is 1. So, this first term goes to 0 and we are left with this minus D_e . So, this is the first term. So, this is equal to minus d , where d is a constant. I will come to it in a bit what these constants mean.

Now, what about V prime at r_e . So, what you do is that they take a derivative of this. Now, clearly this minus 1 does not contribute at all. So, what will have is $2 D_e$ times, you will have exactly this term times e , e raised to r minus r_e . So, this is $a e$ raised to a minus a r minus r_e . So, this is derivative and this has to be evaluated at r equal to r_e and

clearly this is equal to 0. So, when r equal to r_e , this goes to 1. So, 1 minus 1 is 0. So, the first derivative is 0. What about the second derivative?

So, if you want to do the second derivative, then what you get is, you have to take a derivative of this quantity. Now, I will just do this. So, you have $2a$ D_e times the derivative of this whole quantity. Now, what is the derivative of this quantity? So, the first term is just e to the minus a r minus r_e . So, the derivative of that is just minus a e to the minus a r minus r_e , ok.

Now, the second term has e to the minus a r minus r_e into e to the minus a r minus r_e . So, that is I can write it as e to the minus $2a$ r minus r_e and if I take the derivative of that, so I already have an a . So, I have a plus a square e to the minus $2a$ and there will be a factor of 2 coming minus $2a$ r minus r_e and this has to be evaluated at r equal to r_e .

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The image shows a handwritten derivation in a software window titled 'Module 6 - Windows Journal'. The window contains the following mathematical work:

$$V''(r_e) = 2a^2 D_e \left[-ae^{-a(r-r_e)} + 4a^2 e^{-2a(r-r_e)} \right] \Big|_{r=r_e}$$

$$= 2a^2 D_e [2a - 1]$$

SOLUTION (contd.)

$$V'''(r_e) = 2a^2 D_e \left[a e^{-a(r-r_e)} - 4a^2 e^{-2a(r-r_e)} \right] \Big|_{r=r_e}$$

$$= 2a^3 D_e (1 - 4a)$$

The window also shows a toolbar with various drawing tools and a status bar at the bottom with the NPTEL logo and the number 25.

Now, when r equal to r_e , what you see is that this becomes minus a , this becomes $2a$ square. So, this becomes $2a$ square D_e and you have $2a$ minus 1. This is some quantity. So, for completeness let me also do the third derivative V triple prime at r_e . So, this is given by, so I have to take a derivative of this whole thing. So, I have $2a$ square D_e , I have minus e to the minus a . So, that becomes a e to a minus a r minus r_e and then, I have what I had is $2a$. So, plus $2a$ is I have minus $4a$ square e to the minus $2a$ r minus r_e and this whole thing has to be evaluated at r equal to r_e .

So, what you notice is that this works out to be to r equal to r_e . Again these two terms will go away. So, you will get $2a^2 D_e a$ or $2a^3 D_e (1 - 4a)$. So, now I can write my V of r . So, V of r_0 is minus D_e .

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The image shows a handwritten derivation of the potential energy function $V(r)$ for a diatomic molecule. The derivation starts with the third derivative of the potential energy at the equilibrium distance r_e :

$$V'''(r_e) = 2a^2 D_e \left[a e^{-a(r-r_e)} - 4a^2 e^{-2a(r-r_e)} \right] \Big|_{r=r_e}$$

$$= 2a^3 D_e (1 - 4a)$$

Red arrows point to the terms in the second line, labeling them as "quadratic term" and "cubic term". The next line shows the potential energy function $V(r)$ as a sum of three terms:

$$V(r) = -D_e + \frac{(r-r_e)^2}{2!} 2a^2 D_e (2a-1) + \frac{(r-r_e)^3}{3!} 2a^3 D_e (1-4a)$$

Below this, the text "Potential energy of a diatomic molecule" is written. The potential energy function is then given as:

$$V(r) = 2D_e (1 - e^{-a(r-r_e)}) - a e^{-a(r-r_e)}$$

Red underlines are under $2D_e(1 - e^{-a(r-r_e)})$ and $a e^{-a(r-r_e)}$. A graph to the right shows the potential energy curve. The vertical axis is labeled $V(r)$ and the horizontal axis is labeled r . The curve starts at a high positive value for small r , crosses the r -axis at $r = r_e$, reaches a minimum at $r = r_e$ with a value of $-D_e$, and then asymptotically approaches zero as r increases. A vertical dashed line marks $r = r_e$ on the r -axis.

At the bottom left, there is a logo for NPTEL. At the bottom right, the slide number "25" is visible.

Now, the first derivative term is 0. So, the first non-zero term is a second derivative. So, the second derivative term will have a r minus r_e square divided by 2 factorial and you have some quantity which is $2a^2 D_e (2a - 1)$ plus. Now you have r minus r_e whole cube divided by 3 factorial and what you have is $2a^3 D_e (1 - 4a)$. So, this is the final expression for V of r .

Now, a and d are constants. So, what this look like? So, this is actually a potential energy of a typical I will say diatomic molecule. So, this is used to model the potential energy of a diatomic molecule and what does it look like. So, if you look at the graph r s, we said r is greater than or equal to 0. Now, incidentally we see that at r equal to r_e V prime of r_e equal to 0. So, r_e is clearly a minimum. So, V prime of r is equal to let us just write the expression again. So, $2D_e (1 - e^{-a(r-r_e)}) - a e^{-a(r-r_e)}$.

So, you have $2D_e (1 - e^{-a(r-r_e)}) - a e^{-a(r-r_e)}$. Sorry not 2, there should be a^2 here. This is just a . So, what you notice is that this goes to 0 if r equal to r_e . The middle term will go to 0 if r equal to r_e and if r equal to infinity, then you have e raised to minus infinity which goes to 0. So, this term will go to

0 if r equal to infinity. So, this e raise to minus a r minus r_e will go to 0 if r equal to infinity.

So, these are the two places where it goes to 0 and you can look at the second derivative. So, when r equal to r_e , then this has this expression of $2a$ squared D_e times. You have a factor of $2a$ minus 1 and usually $2a$ minus a is chosen to be greater than half. So, this is usually taken as greater than 0. So, usually this is taken to be minimum at r equal to r_e , ok.

So, what this function looks like? So, at r equal to r_e , you have a minimum and what happens is that this function looks like and the value at r equal to r_e is minus D_e . So, this is minus D_e and this is the minimum value of the function and this function essentially looks something like this. So, this is r equal to r_e , it goes to 0 at infinity and also when r equal to 0, when r goes to 0, it actually becomes very large. So, when r_e at this e to the minus e to r_e , some is a fairly large number. So, this is the form of this function and then, this is well known Morse oscillator. So, what is interesting is that if you had only this term which is called the quadratic term and if you only had the quadratic term, then it would look like a harmonic oscillator. So, this is a cubic term. So, this allows you to go beyond the harmonic oscillator, ok.

Now, the next problem that I want to do is problem in statistical mechanics. This is the use of Lagrange's undetermined multipliers, ok.

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Module 6 - Windows Journal

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Problem 2: Lagrange's Undetermined Multipliers

If there are N gas molecules in a system with total energy E ,

N_1 molecules with E_1 ,

N_2 molecules with E_2

\vdots

N_M molecules with E_M

Maximize $W(N_1, N_2, \dots, N_M) = \frac{N!}{N_1! N_2! \dots N_M!}$

Subject to constraints:

$N_1 + N_2 + \dots + N_M = N = \text{constant}$

$N_1 E_1 + N_2 E_2 + \dots + N_M E_M = E = \text{constant}$

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So, here you are told that if there are N gas molecules in a system with total energy E and out of these N 1 molecules have energy E_1 N_2 have molecules have E_2 and so on. N_m have energy E_m , then what you are asked to do is to maximize this w of N_1 up to N_m . This is nothing, but a multinomial function. So, the total number of ways in which you can choose N_1 N_2 up to N_m , such that they all add up to N and the other constraint is that the total energy should be E . So, N_1 times E_1 plus N_2 times E_2 plus up to N_m times. C_m should be some constant.

So, subject to these two constraints you maximize this function. So, let us go ahead, maximize it and then, once we maximize it, we will see some properties of this W . So, let us apply Lagrange's undetermined multipliers. Now, we see this factorial thing appearing.

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SOLUTION: Use $\ln W$ instead of W ; use Stirling's approx.

$$\ln N! = N \ln N - N \quad \text{for large } N$$

$$\ln W = \ln N! - \sum_{i=1}^M \ln N_i!$$

$$= N \ln N - N - \sum_i N_i \ln N_i + \sum_i N_i$$

Apply Lagrange's undetermined multipliers using 2 constraints

$$\frac{\partial \ln W}{\partial N_i} - \lambda \frac{\partial (N_1 + N_2 + \dots + N_M)}{\partial N_i} - \mu \frac{\partial (N_1 E_1 + N_2 E_2 + \dots + N_M E_M)}{\partial N_i} = 0$$

$$-\frac{N_i}{N_i} - \ln N_i + 1 - \lambda - \mu E_i = 0$$

So, I will just emphasize a few things. So, the solution this is done and standard statistical mechanics books is used in of W . So, when W is maximum, natural log of W is also maximum and use stirlings approximation and that stirlings approximation is given by \ln of N factorial equal to $N \ln N$ minus N for large N . So, we just use, these are the tips that you need before you attempt to solve this, ok.

So, what we will do is, \ln of W now natural log of this is \ln of N factorial minus I will just write sum over i equal to 1 to M \ln of N_i factorial and I can write this as natural log of. So, I can write this as $N \ln N$ minus N minus. Now, what I have is here, I have sum

over i of $N_i \ln N_i$ minus sum over i of N_i and what we are going to do is minimize this subject to two constraints. So, now here you have two constraints. So, what we are going to do is when you apply multipliers using two constraints, ok what will say is that do by do N_i of natural log of W minus λ times do by do N_i of constraint 1. So, the first constraint is this N_m , then minus μ times do by do N_i of equal to 0 equal to 0 and you are going to apply this for all i .

So, when you do this, now let us take the derivative of \ln of W with respect to N_i . So, this term will go away, ok. Now, that should be a plus. So, we see that you have N and you have sum over N_i . Now, these two should be equal. So, these two terms can just cancel off. So, then all you have is a derivative with respect to N_i of this quantity. So, let us evaluate this explicitly. So, with respect to N_i of this quantity will give me minus. Now, the term with only the N_i term will appear. So, I will have N_i divided by N_i that is due to derivative of the logarithm and then, minus $\ln N_i$ and I have plus. Now, a derivative with respect to N_i of this is just 1 because this is N_1 plus N_2 . So, the derivative with respect to N_i , there will be one term with just N_i . So, the derivative is just 1. So, this minus λ now do by do N_i of this is just 1. So, then minus μ now do by do N_i of this is just E_i equal to 0, ok.

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Apply Lagrange's undetermined multipliers using 2 constraints

$$\frac{\partial \ln W}{\partial N_i} - \lambda \frac{\partial (N_1 + N_2 + \dots + N_M)}{\partial N_i} - \mu \frac{\partial (N_i E_1 + N_2 E_2 + \dots + N_M E_M)}{\partial N_i} = 0$$

$$-\frac{N_i}{N_i} - \ln N_i + 1 - \lambda - \mu E_i = 0$$

$$-\ln N_i - \lambda - \mu E_i = 0$$

$$\ln N_i = -\lambda - \mu E_i$$

So, finally what you can see is that is that these two terms will cancel and what I have is my final expression. So, let me write this with 0 or I can write N_i is equal to λE_i .

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$$-\ln N_i - \lambda - \mu E_i = 0$$
$$\ln N_i = -\lambda - \mu E_i$$
$$N_i = e^{-\lambda} \cdot e^{-\mu E_i}$$

CONDITION for Maximum W
↓
Most probable distribution

From Statistical Mechanics, we can show $\mu = \frac{1}{k_B T}$; $e^{-\lambda} = \frac{N}{\sum_{i=1}^M e^{-\mu E_i}}$

$$\sum_{i=1}^M N_i = N = e^{-\lambda} \sum_{i=1}^M e^{-\mu E_i}$$

So, this is $\ln N_i$ and you get a very useful relation. Now, if I write it in terms of exponential, I get N_i is equal to $e^{-\lambda}$ times $e^{-\mu E_i}$. This is the solution, ok.

Now, this is what we are interested in. So, this is the condition for maximum W . So, it is also in statistical mechanics and this is called the most probable distribution. Now, λ and μ are undetermined multipliers. So, again you know from statistical mechanics, we can show μ is equal to $1/k_B T$ and $e^{-\lambda}$ is just N over the sum of i equal to 1 to M of $e^{-\mu E_i}$ because you can show this equal to N over.

So, you can show this by summing both sides. So, basically sum over i equal to 1 to M of N_i is equal to N is equal to $e^{-\lambda}$ sum over i equal to 1 to M of $e^{-\mu E_i}$. So, you can get this at basically once you show these two, then you get a very well known expression.

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The image shows a whiteboard with handwritten mathematical expressions and text. At the top left, a boxed equation reads $\frac{N_i}{N} = \frac{e^{-\mu E_i}}{\sum_{j=1}^M e^{-\mu E_j}}$. To its right, the text "CANONICAL ENSEMBLE DISTRIBUTION FUNCTION" is written, with a circled equation $\mu = \frac{1}{k_B T}$ below it. Arrows point from the text "Boltzmann constant" and "Temperature" to the terms k_B and T in the circled equation, respectively. Below the boxed equation, another equation reads $\frac{N_i}{N} = \frac{1}{Q} \cdot e^{-\frac{E_i}{k_B T}}$. To the right of this, the equation $Q = \sum_{j=1}^M e^{-\frac{E_j}{k_B T}}$ is written, with the text "Partition Function" underneath it. The whiteboard also features a toolbar at the top and an NPTEL logo at the bottom left.

So, that is N_i by N is equal to e to the minus μE_i divided by sum over i equal to 1 to M or let me call this j equal to 1 to M just to distinguish it from i that we are using e to the minus μE_j . So, this is called the canonical ensemble distribution function and if you use $\mu = 1/k_B T$ I should emphasize. So, if μ is equal to $1/k_B T$ which is called Boltzmann constant and T is temperature, ok.

So, this is a very well known expression in statistical mechanics. So, sometimes this denominator is called the partition function Q and you have e to the minus $E_i/k_B T$, where Q is equal to sum over j equal to 1 to M e to the minus $E_j/k_B T$ is sum over all states. So, this is called the partition function.

So, what we have seen is that the Lagrange's method of undetermined multipliers that we have used this method in order to calculate the distribution of a gas when at a given temperature. Now, I did not show one part. I did not show how you get $\mu = 1/k_B T$. So, this is actually slightly more involved. I mean this you need to use some more thermodynamics. You cannot just do it by mathematics. So, the connection between μ and temperature is caught from thermodynamics, but essentially you get the basic form, you get this expression just from by using Lagrange's method of undetermined multipliers and maximizing this quantity W . So, we identified that W is the quantity that we have to maximize. So, this quantity W is what we have to maximize and we took logarithm of W and we use Stirling's approximation which is extremely valid. If you have

a large ends which is what happens when you have a very large number of molecules. So, statistical mechanics whole when the number of molecules is of the order of avogadros number, ok.

So, this stirlings approximation is almost exact. So, with this I will conclude module 6 and in the next module, we will start looking at qualitative methods for non-linear differential equations.

Thank you.