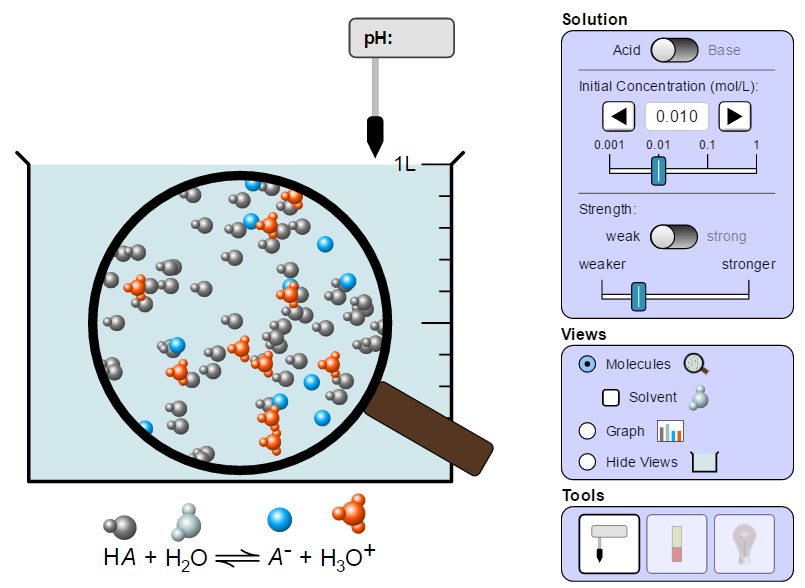
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**Activity: Strong vs. Weak Acids** (Targets 11.3, 11.7, 11.8)



**Instructions:**

* Go to the following simulation: <https://phet.colorado.edu/sims/html/acid-base-solutions/latest/acid-base-solutions_en.html>
* Click on “My Solution”

**Pre-lab:**

1. Write the dissociations for the follow acids:

HCl (strong): \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

HC2H3O2 (weak): \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. What is the difference in how strong and weak acids dissociate in water?

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| **Graph for Strong Acid** |

**Part A: Strong Acid dissociations versus Weak Acid Dissociations**

1. In the simulation, use the slider and set the initial concentration to 0.010 M and set the acid strength to “strong”.
2. Under Views, click molecules. How do the numbers of molecules of the molecular acid compare to the number of cations and anions?

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1. Under views, click Graph. Sketch the bar graph in the box provided. Ignore water.

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| **Graph for Weak Acid** |

1. Under view click back to “molecules”. Now change the acid strength to “weak”. How does the weak acid bar graph differ from the strong acid? Why?

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1. Under views, click Graph. Sketch the bar graph to the right. Ignore water.
2. Summary Questions: Which has more [H+] in solution – the strong or weak acid? How do you know from looking at pH that this solution has more [H+]? (Hint: Is the solution more acidic or less acidic?)

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**Part B: pH, [H+] for strong acids versus weak acids.**

1. In the simulation, set initial concentration to 0.0010 M and set the strength to “strong”. Under views, click “molecules”.
2. Move the pH probe into the beaker. Record the pH in the table below.
3. Change the strength to “weak”. Record the pH in the table below.
4. Repeat for the other concentrations listed in the table below.

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| --- | --- | --- |
| Initial Concentration | pH for Strong Acid | pH for Weak Acid |
| 0.0010 M |  |  |
| 0.010 M |  |  |
| 0.10 M |  |  |
| 1.0 M |  |  |

1. In general, what happens to pH as acid concentration increases?

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1. Compare the strong acid and the weak acid at the same concentration.
   1. How do their pHs differ?

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* 1. Explain why the pH for a strong acid would be different from a weak acid. Relate how they dissociate, their [H+] and their pHs.

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